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First-principles study of localized and delocalized electronic states in crystallographic shear phases of niobium oxide

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Crystallographic shear phases of niobium oxide form an interesting family of compounds that have received attention both for their unusual electronic and magnetic properties, as well as their performance as intercalation electrode materials for lithium-ion batteries. Here we present a first-principles density-functional theory study of the electronic structure and magnetism of $\text{H-Nb}_2\text{O}_5$, $\text{Nb}_{25}\text{O}_{62}$, $\text{Nb}_{47}\text{O}_{116}$, $\text{Nb}_{22}\text{O}_{54}$, and $\text{Nb}_{12}\text{O}_{29}$. These compounds feature blocks of niobium-oxygen octahedra as structural units, and we show that this block structure leads to a coexistence of flat and dispersive energy bands, corresponding to localized and delocalized electronic states. Electrons localize in orbitals spanning multiple niobium sites in the plane of the blocks. Localized and delocalized electronic states are both effectively one-dimensional and are partitioned between different types of niobium sites. Flat bands associated with localized electrons are present even at the GGA level, but a correct description of the localization requires the use of GGA + U or hybrid functionals. We discuss the experimentally observed electrical and magnetic properties of niobium suboxides in light of our results, and argue that their behavior is similar to that of n -doped semiconductors, but with a limited capacity for localized electrons. When a threshold of one electron per block is exceeded, metallic electrons are added to existing localized electrons. We propose that this behavior of shear phases is general for any type of n -doping, and should transfer to doping by alkali metal (lithium) ions during operation of niobium oxide-based battery electrodes. Future directions for theory and experiment on mixed-metal shear phases are suggested.

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I. INTRODUCTION

Transition metal oxides form a fascinating class of compounds with interesting electronic, magnetic, and crystallographic structures. The phase diagram of niobium oxide is especially rich, with a large number of reported phases for Nb_2O_5 [1–3], in addition to NbO and NbO_2 . The high-temperature Nb_2O_5 polymorph ($\text{H-Nb}_2\text{O}_5$) can be regarded as the parent compound of a family known as crystallographic shear (or Wadsley-Roth) phases [4,5]. In these phases, niobium is present in octahedral coordination, but the Nb/O ratio of Nb_2O_5 prevents the formation of purely corner-sharing octahedra. Instead, the structure must include some amount of edge-sharing connections between octahedra. The crystal structures of these compounds consequently consist of *blocks* of corner-sharing octahedra of size $n \times m$ that are connected to neighboring blocks via crystallographic shear planes of edge-sharing connections. In the direction perpendicular to the $n \times m$ plane the units connect infinitely, and tetrahedrally coordinated ions are sometimes present to fill voids in the structure. By reduction of Nb_2O_5 , small amounts of Nb^{4+} can be incorporated, and a series of $\text{Nb}_2\text{O}_{5-\delta}$ compounds form. These suboxides include $\text{Nb}_{25}\text{O}_{62}$, $\text{Nb}_{47}\text{O}_{116}$, $\text{Nb}_{22}\text{O}_{54}$, and

two polymorphs of $\text{Nb}_{12}\text{O}_{29}$ with different crystal symmetries (Fig. 1, Table I). The metal-oxygen octahedra in these compounds are strongly distorted due to a combination of electrostatic repulsion between transition metal ions and the second-order Jahn-Teller effect [6,7]. Niobium sites in the center of the block are less distorted than those at the periphery. The structural principle of blocks as the building unit, as introduced by Wadsley and Roth, also applies to phases in the TiO_2 - Nb_2O_5 and WO_3 - Nb_2O_5 phase diagrams [5].

The fully oxidized parent compound Nb_2O_5 is a wide band gap insulator. Low concentrations of valence electrons are introduced through n -type doping to form the $\text{Nb}_2\text{O}_{5-\delta}$ phases. This reduction changes the crystal structure, but the structural motif of the blocks is retained, which makes the niobium suboxides an excellent series of phases to study the interplay between charge state and crystal structure. Magnetic susceptibility measurements show that all $\text{Nb}_2\text{O}_{5-\delta}$ phases are paramagnetic, with the number of localized moments increasing with δ [8,9]. Spin interactions are antiferromagnetic and their strength increases with the level of reduction, as indicated by their Curie-Weiss constants. However, only the monoclinic $\text{Nb}_{12}\text{O}_{29}$ phase is found to exhibit long-range antiferromagnetic order, with an ordering temperature of 12 K [10,11]. Electrical conductivity measurements show that all $\text{Nb}_2\text{O}_{5-\delta}$ phases exhibit thermally activated transport, except for $\text{Nb}_{12}\text{O}_{29}$, which is metallic down to 2 K [8,12]. Both electrical and optical measurements indicate that the

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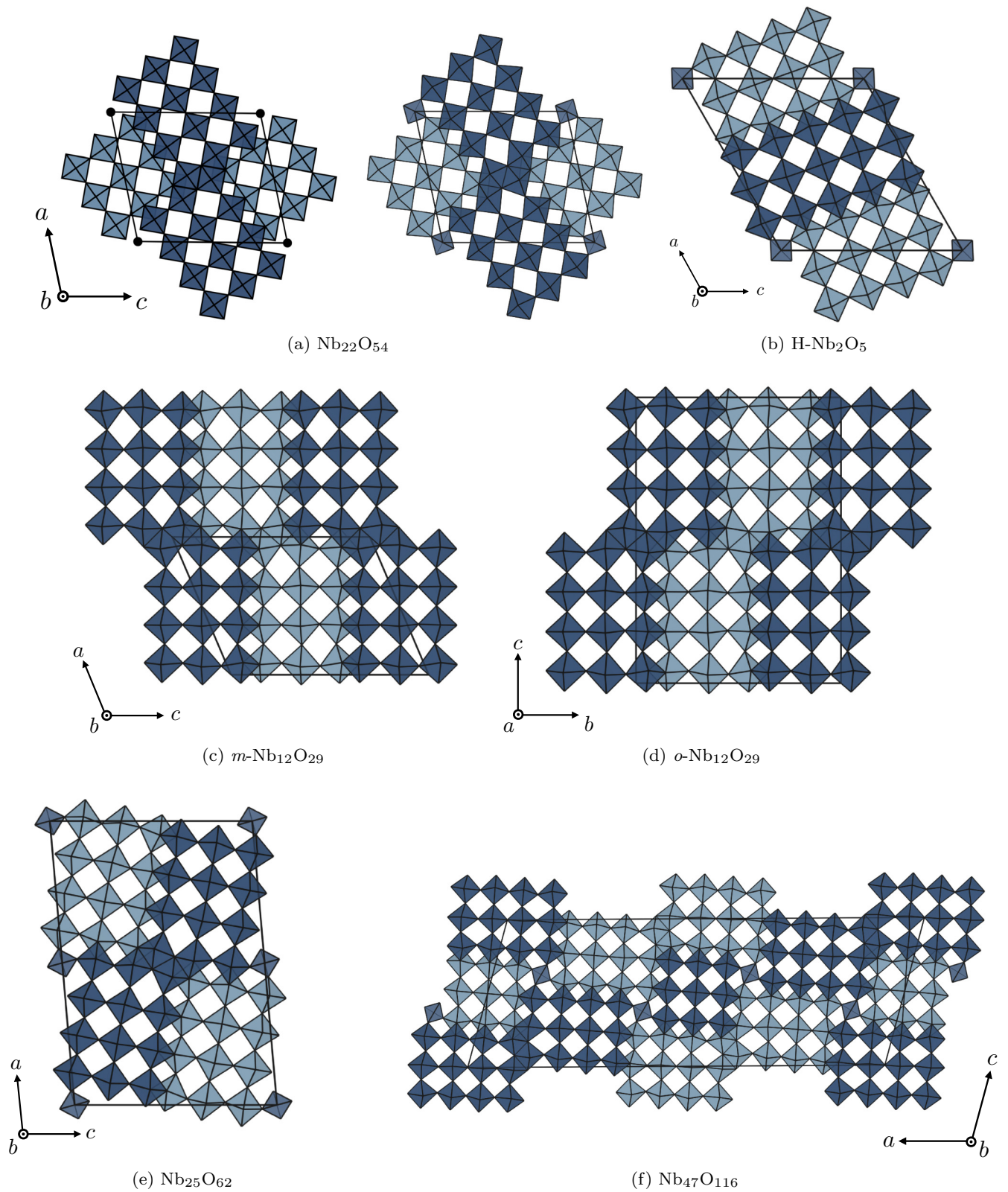


FIG. 1. (a) Idealized (left) and locally distorted (right) crystal structure of $\text{Nb}_{22}\text{O}_{54}$. The tetrahedral site is shown as a black dot in the idealized structure. Crystal structures of (b) $\text{H-Nb}_2\text{O}_5$, (c) monoclinic $\text{Nb}_{12}\text{O}_{29}$, (d) orthorhombic $\text{Nb}_{12}\text{O}_{29}$, (e) $\text{Nb}_{25}\text{O}_{62}$, and (f) $\text{Nb}_{47}\text{O}_{116}$. Light and dark colored blocks are offset by half of the lattice parameter perpendicular to the plane of the page. Unit cells are outlined in black.

TABLE I. Structural properties of niobium (sub)oxides. Experimental and DFT optimized lattice parameters a , b , and c are given in Å. Structural optimizations with DFT + U were performed with a U_{eff} value of 4.0 eV on niobium d orbitals. N_{e^-} denotes number of electrons introduced by doping. Difference between experimental and DFT space group choices related to ordering of tetrahedral sites (see text).

Compound	O/Nb	N_{e^-} /block	Space group	Source	a	b	c	β	Block size(s)
H-Nb ₂ O ₅	2.500	0	$P2/m$	Expt. [2]	21.153	3.823	19.356	119.80°	$3 \times 4, 3 \times 5$
			$P2$	PBE	21.433	3.841	19.614	119.85°	
Nb ₂₅ O ₆₂	2.480	1/2	$C2$	Expt. [8]	29.78	3.825	21.14	94.7°	3×4
				PBE	30.224	3.84	21.44	95.0°	
Nb ₄₇ O ₁₁₆	2.468	2/3	$C2$	Expt. [8]	57.74	3.823	21.18	105.3°	$3 \times 3, 3 \times 4$
				PBE	58.43	3.84	21.44	105.3°	
Nb ₂₂ O ₅₄	2.455	1	$P2/m$	Expt. [14]	15.749	3.824	17.852	102.029°	$3 \times 3, 3 \times 4$
			$P2$	PBE	15.931	3.842	18.036	102.06°	
			$P2$	PBE + U	15.935	3.836	18.061	101.99°	
m -Nb ₁₂ O ₂₉	2.417	2	$A2/m$	Expt. [21]	15.695	3.831	20.723	113.103°	3×4
				PBE	15.903	3.832	20.967	113.1°	
				PBE + U	15.885	3.837	20.950	113.09°	
o -Nb ₁₂ O ₂₉	2.417	2	$Cmcm$	Expt. [14]	3.832	20.740	28.890	–	3×4
				PBE	3.833	20.955	29.241	–	
				PBE + U	3.836	20.961	29.204	–	

electron transport in the Nb₂O_{5- δ} phases is effectively one-dimensional along the block columns [13]. Despite the evidence for localized electrons, single crystal x-ray diffraction studies on Nb₂₂O₅₄ and o -Nb₁₂O₂₉ have not found evidence for charge ordering [14]. Additional studies have been performed on Nb₁₂O₂₉ indicating the presence of localized as well as delocalized electrons [12,15,16].

Despite their interesting physical properties, the niobium suboxides have not been investigated with first-principles methods, and the relationship between the level of reduction and the electronic and magnetic properties remains unclear. Two previous first-principles studies have examined the two polymorphs of Nb₁₂O₂₉, with rather different conclusions regarding the electronic structure, despite their very strong structural similarity [17,18]. Additionally, there is new interest in crystallographic shear phases due to their excellent performance as electrode materials in batteries [19,20], and the electronic structure of the suboxides is likely to be similar to that of other shear phases.

In this article we study six different niobium (sub)oxides using density-functional theory calculations and establish common principles governing their electronic and magnetic behavior. As the $4d$ band in these materials is highly complex, we first examine Nb₂₂O₅₄ in detail, and then present results on Nb₁₂O₂₉, Nb₄₇O₁₁₆, Nb₂₅O₆₂, and H-Nb₂O₅. We show that all of these structures host flat and dispersive bands, which correspond to localized and delocalized electronic states. Each block can host a single localized state in the block plane that is spread over multiple niobium sites. Delocalized states are present along the shear planes. These results are independent of the n -type doping, and alkali metal doped (lithiated) shear phases show similar features to the suboxides, which has implications for their use as battery electrodes. We discuss experimental studies of electrical and magnetic properties of the suboxides in terms of a consecutive filling of localized and delocalized states. Finally, based on these results, we suggest new directions for theory and experiment.

II. METHODS

All density-functional theory calculations were performed using the plane-wave DFT code CASTEP [22] (version 18.1). Pseudopotentials including Nb $4s$, $4p$, $4d$, and $5s$, O $2s$ and $2p$, and Li $1s$ and $2s$ states were used for all calculations. Calculations using hybrid functionals employed norm-conserving pseudopotentials [23], all other calculations were performed using Vanderbilt ultrasoft pseudopotentials [24]. Crystal structures were obtained from the Inorganic Crystal Structure Database (ICSD) [25] when available. The structure of Nb₄₇O₁₁₆ was constructed as described in Ref. [8] as a unit cell intergrowth of Nb₂₅O₆₂ and Nb₂₂O₅₄ since no crystallographic data, other than the lattice parameters, was available. The space groups of H-Nb₂O₅ and Nb₂₂O₅₄ are reported as both $P2$ and $P2/m$ in the literature [2,8,14]. These two space group choices differ only in the full or partial occupancy of the tetrahedral site. For modeling purposes, the ion on the tetrahedral site has to be ordered, resulting in space group $P2$. Atomic positions and lattice parameters of the structures were relaxed using the gradient-corrected Perdew-Burke-Ernzerhof (PBE) functional [26], until the maximum force on any atom was smaller than 0.01 eV/Å. The calculations used a plane-wave kinetic energy cutoff of 800 eV for ultrasoft pseudopotentials, and 900 eV for norm-conserving pseudopotentials, unless otherwise stated. The Brillouin zone was sampled with a Monkhorst-Pack grid [27] finer than $2\pi \times 0.03 \text{ \AA}^{-1}$. Lattice parameters obtained from the structural relaxations are listed in Table I, and agree very well with the experimental values. Crystallographic information files (CIF) of the PBE optimized structures are available in the Supplemental Material [28]. All electronic structure calculations were performed for antiferromagnetic spin arrangements in the conventional unit cells, as antiferromagnetic spin interactions are observed experimentally [8].

Semilocal density functionals suffer from self-interaction error, which can be alleviated by the use of DFT + U . For

calculations in this work, the DFT + U implementation in CASTEP [29] was used, which defines an effective U value $U_{\text{eff}} = U - J$. A value of $U_{\text{eff}} = 4.0$ eV was chosen for the Nb d orbitals, in line with other studies on niobium oxides that employed similar implementations of DFT + U within plane-wave codes [30]. The results presented herein are mostly insensitive to the exact value of the U_{eff} parameter if it lies in the range 3–5 eV, even though the value of the band gap does depend on the choice of the U_{eff} parameter. The structures of Nb₂₂O₅₄ and the Nb₁₂O₂₉ polymorphs were additionally optimized with PBE + U , and the results are listed in Table I. PBE and PBE + U lattice parameters agree closely, and PBE + U band structure and density of states calculations for compounds other than Nb₂₂O₅₄ and Nb₁₂O₂₉ were performed on PBE optimized structures.

Hybrid functionals are another way to correct the self-interaction error of semilocal functionals. The range-separated HSE06 functional [31] was used to calculate the band structure for Nb₂₂O₅₄. Due to the significant additional expense incurred by the use of hybrid functionals, the computational parameters for the calculations of band structures at the HSE06 level are coarser. The unit cell of Nb₂₂O₅₄ contains 610 valence electrons, but since the cell is rather short in one particular dimension and extended in the other two, one cannot use only the Γ point in the Brillouin zone (BZ) sampling. Instead, a $1 \times 5 \times 1$ Γ -centered \mathbf{k} -point grid was used in the HSE06 self-consistent field calculations for Nb₂₂O₅₄.

Band structure calculations were performed for high-symmetry Brillouin zone directions according to those obtained from the SeeK-path package [32], which relies on spglib [33]. A spacing between \mathbf{k} -points of $2\pi \times 0.025 \text{ \AA}^{-1}$ was used. Density of states calculations were performed with a grid spacing of $2\pi \times 0.01 \text{ \AA}^{-1}$, and the results were post-processed with the OptaDOS package [34], using the linear extrapolative scheme [35,36]. The c2x [37] utility and VESTA [38] were used for visualization of wave function and density data. Data analysis and visualization was performed with the matador [39] package.

III. RESULTS

A. Nb₂₂O₅₄

Nb₂₂O₅₄ crystallizes in space group $P2/m$ [14], and shows an ordered mixture of 3×3 and 3×4 blocks of octahedra, in addition to a tetrahedral site [Fig. 1(a)]. Assuming an ionic model, the compound can be described as $(\text{Nb}^{5+})_{20}(\text{Nb}^{4+})_2(\text{O}^{2-})_{54}$, with two $4d$ electrons per 22 Nb atoms ($1 e^-$ per block, Table I).

The \mathbf{a} and \mathbf{c} lattice vectors of Nb₂₂O₅₄ are longer than \mathbf{b} , which is perpendicular to the block plane (Fig. 1, Table I). The Brillouin zone (BZ) therefore has one long (along \mathbf{b}^*) and two short directions. The PBE + U spin-polarized band structure and electronic density of states (DOS) of Nb₂₂O₅₄ show a large gap between the valence and conduction bands, which are of oxygen $2p$ and niobium $4d$ character, respectively (Fig. 2). Two fully occupied flat bands (one for each spin) lie within the band gap, leading to the peaks in the DOS below the Fermi level. The flat bands have a very small one-dimensional

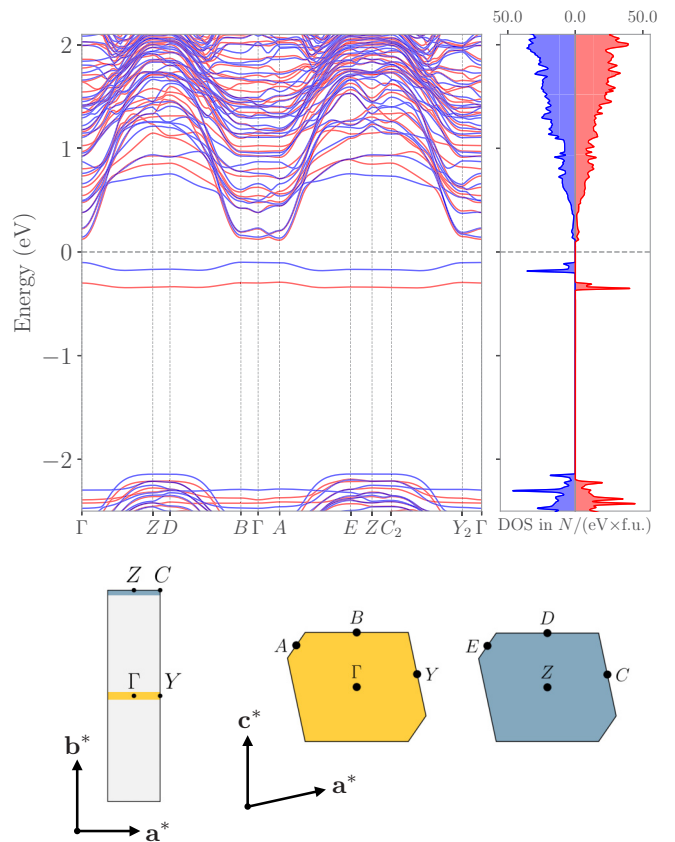


FIG. 2. Spin-polarized band structure and electronic density of states of Nb₂₂O₅₄ (PBE + U , $U_{\text{eff}} = 4$ eV). Up and down spins colored in red and blue. High symmetry points are marked on slices through the first Brillouin zone. The flat bands below the Fermi level (dashed line) represent localized states.

dispersion, as evidenced by the shapes of the corresponding peaks in the DOS, and represent localized states. In addition to the flat bands, a set of dispersive bands exists just above the Fermi level, which show the largest dispersion along \mathbf{b}^* . The separation between the flat bands and the rest of the conduction states is smallest at special points lying in the \mathbf{a}^* - \mathbf{c}^* plane of Γ (Y , A , B), and largest in the parallel plane at the BZ boundary (Z , C , D , E). Due to this pattern, the dispersive bands are also effectively one-dimensional.

With 12 inequivalent niobium sites in the unit cell of Nb₂₂O₅₄, site-resolved projected densities of states (PDOS) are complicated and difficult to interpret. More insight is gained by summing PDOS for sets of sites. Figure 3(a) shows the projection onto different types of niobium sites within the structure, which are classified as central and peripheral, depending on where they sit within the block. We note two things: (1) Both peripheral and central niobium sites contribute to the localized states, even though the contribution of the central sites is greater given the ratio of the two; and (2) only peripheral niobium sites contribute to the unoccupied density of states above the Fermi level (until 0.5 eV above), the contribution from the central sites is exactly zero. The PDOS resolved by block in Fig. 3(b) demonstrates that one localized state is contained in the 3×4 block, and the other, lower energy one, in the 3×3 block. Both blocks contribute

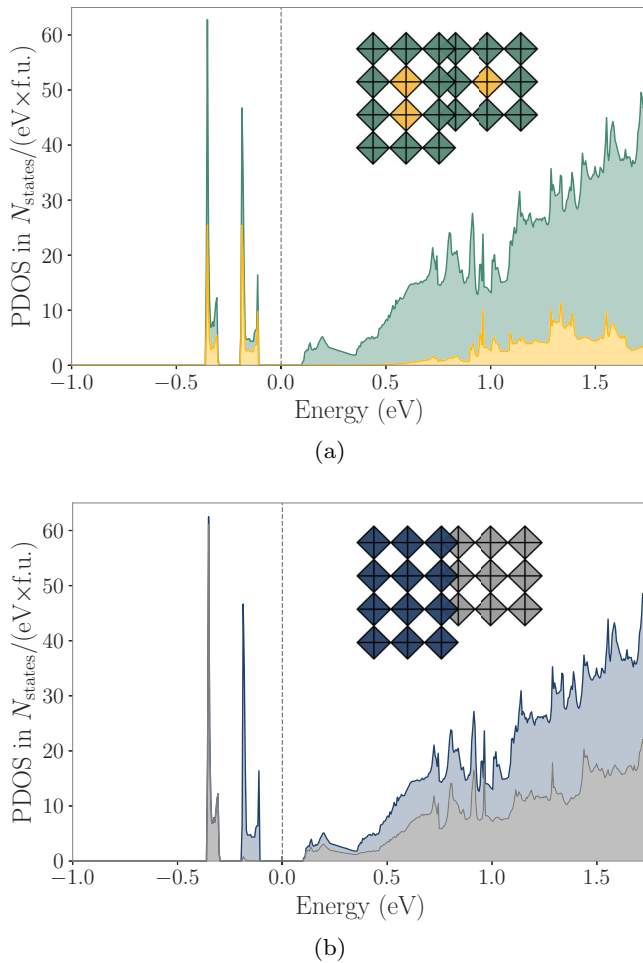


FIG. 3. Spin-summed projected density of states (PBE + U) for $\text{Nb}_{22}\text{O}_{54}$. Fermi level is indicated by the dashed line. (a) PDOS for central (gold) and peripheral (green) niobium sites. (b) DOS projected for sites in different blocks, demonstrating separate localization of electrons in 3×3 and 3×4 blocks. Contributions from sites are proportional to the shaded area.

roughly equally to the density of unoccupied conduction states.

Spin density in $\text{Nb}_{22}\text{O}_{54}$ is predominantly located on the central niobium sites [Fig. 4(a)], which also dominate the relevant states as seen from the PDOS [Fig. 3(a)]. One spin is located in each block, and the spin arrangement is antiferromagnetic between the two blocks. However, the ferromagnetic arrangement is only marginally higher in energy (less than 1 meV), indicating very weak spin interactions that are likely a result of the long (nm) magnetic interaction lengths. Kohn-Sham orbitals that are occupied by these localized electrons span the entire block, but only have contribution from niobium sites in the same block [Fig. 4(b)]. The flat dispersion is a result of the very weak face-on overlap (δ -overlap) between these orbitals along **b**. Both localized orbitals are similar in appearance, despite the different sizes of the blocks. This suggests that the presence of these states is a general feature of block-type structures.

The results presented above were obtained from PBE+ U ($U_{\text{eff}} = 4$ eV) calculations. The $\Gamma \rightarrow Z$ segment of the

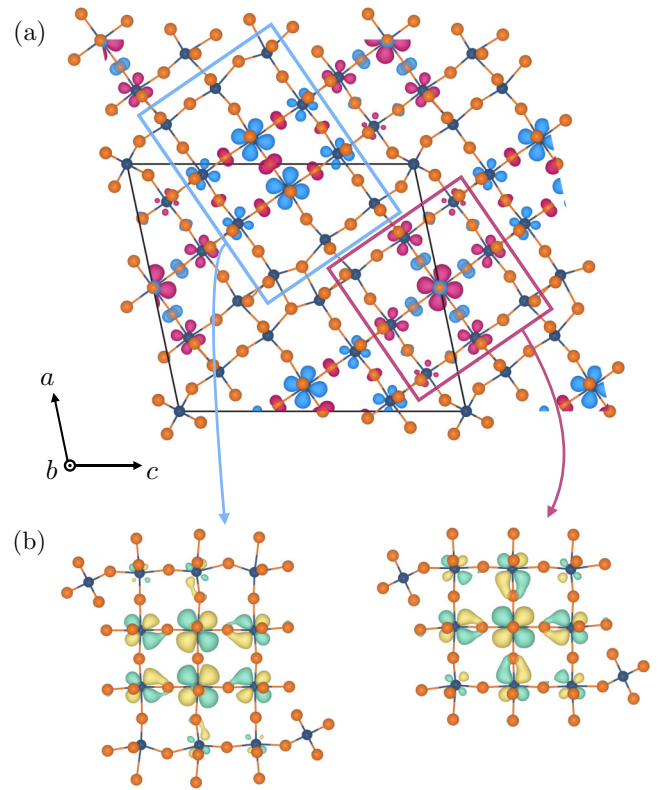


FIG. 4. (a) Spin density plot of $\text{Nb}_{22}\text{O}_{54}$. Niobium and oxygen shown in dark blue and orange, respectively. Purple and light blue represent up and down spin density, respectively. The rectangles outline the 3×4 and 3×3 blocks. Spin density isosurface drawn at a value of $0.03 e^-/\text{\AA}^3$. (b) Kohn-Sham orbitals associated with localized states (flat bands in Fig. 2) in 3×4 and 3×3 blocks, different phases of the orbitals shown in yellow and light green.

$\text{Nb}_{22}\text{O}_{54}$ band structure computed with HSE06 and PBE is compared to the PBE + U result in Fig. 5. Only the up-spin component is shown, which is associated with the localized electron in the 3×3 block. The band structure looks similar

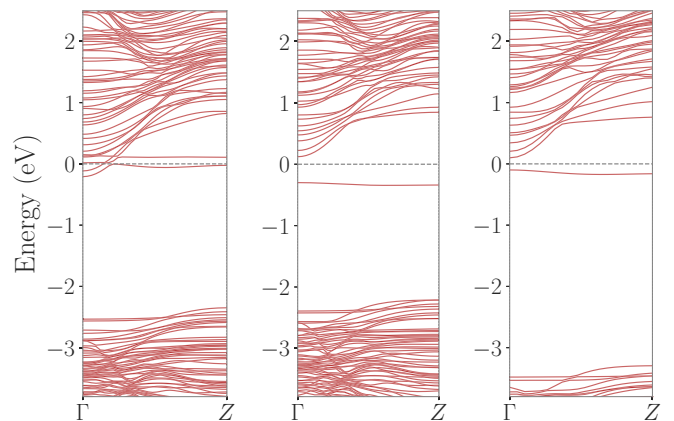


FIG. 5. $\Gamma \rightarrow Z$ segment of the band structure of $\text{Nb}_{22}\text{O}_{54}$ calculated with different levels of theory. PBE, PBE + U ($U_{\text{eff}} = 4.0$ eV), and HSE06, from left to right. Only one spin component is shown for clarity.

for all functionals, and importantly, the relevant feature of localized states, i.e. the flat bands, are present even at the PBE level. However, PBE places the flat bands within the dispersive conduction bands, and both are partially occupied, while both DFT + U and the HSE06 functional place the flat bands below the other conduction bands. PBE also places the opposite spin partner of the localized state in the other block much lower in energy than either PBE + U or HSE06. The precise placement of the flat bands depends on the U value, but in the tested range of 2–5 eV the flat bands are placed below the conduction bands, and the gap between them increases by approximately 200 meV per increase in U by 1 eV. The degree of electron localization depends on the presence of a gap between flat and dispersive bands. PBE implies metallic behavior with localized electrons, while HSE06 and PBE + U show full localization of the electrons. A major difference between the HSE06 and PBE or PBE + U calculations is the size of the gap between valence and conduction bands, which is larger by approximately 1.2 eV for HSE06 compared to PBE + U . The spin density and Kohn-Sham orbitals were plotted from the output of PBE + U calculations, but we note that the results from PBE and HSE06 are visually indistinguishable from the PBE + U results.

B. Nb₁₂O₂₉

Nb₁₂O₂₉ is more reduced than Nb₂₂O₅₄ and hosts two 4*d* electrons per 12 niobium sites (i.e., 2 per block, Table I). The two Nb₁₂O₂₉ polymorphs are structurally similar, and only differ in the long-range arrangement of the blocks; in the monoclinic polymorph the blocks form a ribbon along a , while in the orthorhombic structure the blocks zigzag along c (Fig. 1).

The band structure of monoclinic Nb₁₂O₂₉ shows two flat bands (one for each spin), which lead to two peaks in the DOS (Fig. 6). The shape of the real-space unit cell results

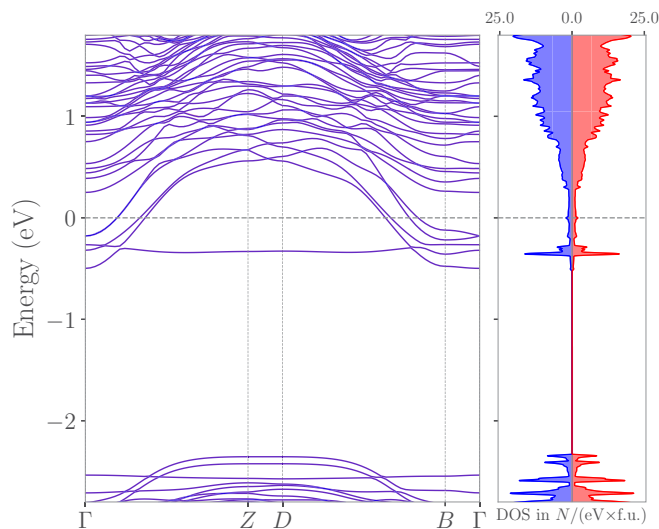


FIG. 6. Band structure and density of states of monoclinic Nb₁₂O₂₉ (PBE + U). Fermi level indicated by a dashed line. Up and down spins colored in red and blue, respectively. Flat and dispersive bands are present, with strong similarity to those in Nb₂₂O₅₄.

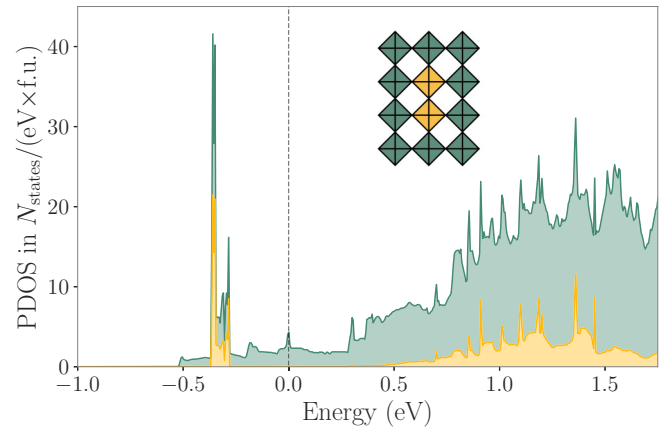


FIG. 7. Projected density of states (PBE + U) for central (gold) and peripheral (green) niobium sites in monoclinic Nb₁₂O₂₉. Central niobium sites contribute to the density of occupied states only in a narrow window that is associated with the flat bands.

in a Brillouin zone with two short and one long dimension, and the band structure path segments are similar to those in Nb₂₂O₅₄. The bands for both spins lie exactly on top of each other due to the symmetry of the crystal structure, even though there is a spatial separation of spins [Fig. 8(a)]. The flat bands coexist with more dispersive conduction bands, which show a dispersion which is largest in the \mathbf{b}^* direction, making them effectively one-dimensional. Independent of the position of the flat bands, the larger number of electrons per block requires that some of the electrons fill dispersive conduction bands. This indicates a structural capacity for localized electrons. In Nb₁₂O₂₉ flat and dispersive bands are interspersed, while in Nb₂₂O₅₄, the flat bands lie below the rest of the d bands (cf. Fig. 2). Similar to Nb₂₂O₅₄, the central Nb sites contribute exclusively to the occupied density of states in a narrow region that is associated with the flat bands (Fig. 7). The remainder of the conduction states involve contributions from the peripheral sites.

Kohn-Sham orbitals of the localized states [Fig. 8(c)] look remarkably similar to those in Nb₂₂O₅₄, and are predominantly made up of Nb d orbitals lying within the plane of the block. Electrons occupying these localized states are responsible for the nonzero spin density [Fig. 8(a)]. Orbitals associated with dispersive bands [Figs. 8(b) and 8(d)] are made up of d_{xy} and d_{yz} atomic orbitals that are parallel to the crystallographic shear planes. The band dispersion along $\Gamma \rightarrow Z$ is explained by a reduction of in-phase overlap of the constituent atomic orbitals along the real-space \mathbf{b} direction. The fact that the contributing atomic orbitals are parallel to the crystallographic shear planes and overlap face-on [Figs. 8(b) and 8(d)], rather than end-on, can be understood from a crystal field argument. For a transition metal ion in an ideal octahedral crystal field, the t_{2g} orbitals form a degenerate set. The MO₆ octahedra in shear phases, however, are far from ideal. When the degeneracy of the t_{2g} orbitals is lifted by a distortion, those d orbitals that do not overlap with any σ -type ligand orbitals will be lowest in energy and contribute to the low-energy d bands.

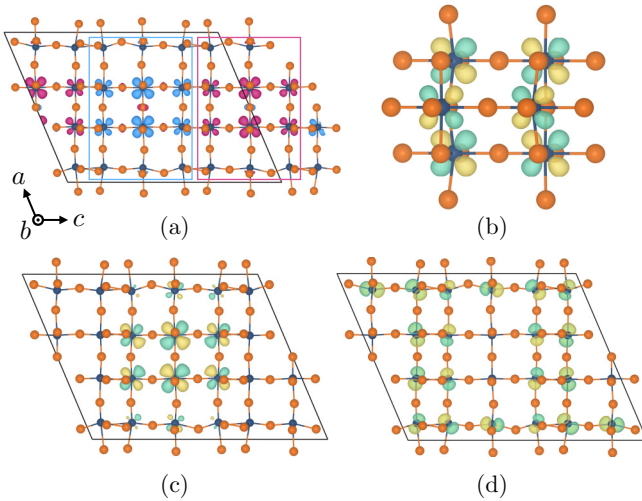


FIG. 8. Spin density plot (a) and orbitals associated with localized (c) and delocalized (b) and (d) states in monoclinic $\text{Nb}_{12}\text{O}_{29}$. Spin density (a) is predominantly located on the central niobium sites, and results from the occupation of localized states (c). Delocalized states have no contribution from the central niobium sites.

The band structure of the orthorhombic $\text{Nb}_{12}\text{O}_{29}$ polymorph (Fig. 9) looks very similar to that of the monoclinic polymorph. In fact, the results on the monoclinic polymorph presented above are transferable to the orthorhombic one. Experimentally, both polymorphs are metallic and show antiferromagnetic spin interactions, but only the monoclinic polymorph shows long-range magnetic order [11]. The magnetic susceptibility of $m\text{-Nb}_{12}\text{O}_{29}$ can be fit using the Bonner-Fisher form, possibly indicating one-dimensional magnetism [15]. The differences are clearly subtle, and the small energy differences (10–20 K, around 1 meV) make comparisons using density-functional theory total energy differences difficult.

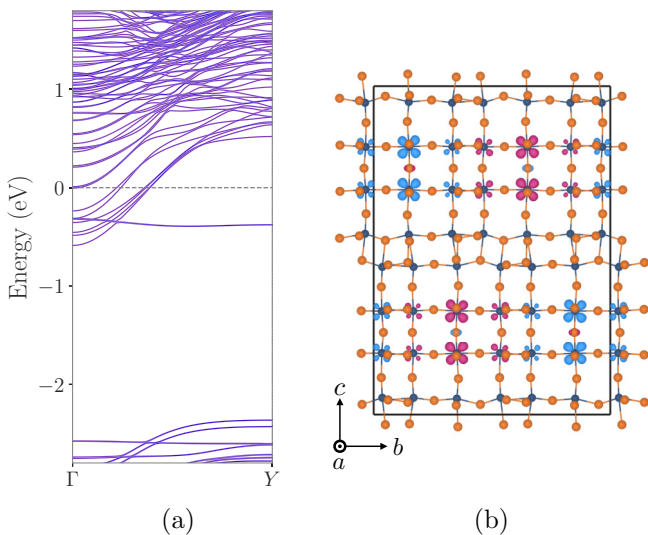


FIG. 9. Band structure (PBE + U) (a) and spin density (b) of orthorhombic $\text{Nb}_{12}\text{O}_{29}$. $Y = \mathbf{a}^*/2$. The orthorhombic and monoclinic $\text{Nb}_{12}\text{O}_{29}$ polymorphs show a strong similarity in their band structure and spin density distribution [cf. Figs. 6 and 8(a)].

However, the picture of the electronic structure of $\text{Nb}_{12}\text{O}_{29}$ that emerges is clear: for both polymorphs, conductivity and local moment magnetism are provided by different sets of electrons. Our conclusions on the orthorhombic polymorph are broadly in line with the first-principles study of Lee and Pickett [18]. Those authors also found a coexistence of localized and delocalized electrons, with the localized spin residing in a large orbital dominated by the central niobium sites of the blocks, with delocalized electrons forming another subset. Our results as well as experimental studies using heat capacity measurements [16] and μSR spectroscopy [15] establish the presence of localized magnetic electrons in $m\text{-Nb}_{12}\text{O}_{29}$. We note that a previous study suggested the presence of itinerant moments in $m\text{-Nb}_{12}\text{O}_{29}$ on the basis of GGA calculations [17]. However, the high density of states at the Fermi level that was described to be the reason for the itinerant magnetism in fact arises from the flat band representing a localized state.

C. $\text{Nb}_{25}\text{O}_{62}$ and $\text{Nb}_{47}\text{O}_{116}$

The compounds $\text{Nb}_{25}\text{O}_{62}$ and $\text{Nb}_{47}\text{O}_{116}$ are less reduced than $\text{Nb}_{22}\text{O}_{54}$ and host less than one electron per structural block unit (Fig. 1, Table I). The mutually occurring localized and delocalized electronic states that were found above for $\text{Nb}_{22}\text{O}_{54}$ and monoclinic $\text{Nb}_{12}\text{O}_{29}$ are also present in $\text{Nb}_{25}\text{O}_{62}$ and $\text{Nb}_{47}\text{O}_{116}$. Localized states in blocks of the same size are nearly degenerate, and since only a fraction of the localized states is occupied (less than 1 electron per block), it is very difficult in a first-principles calculation to localize the electrons within a specific block. This could be done if the occupation of particular bands was constrained. Similarly, since the energy of the localized states depends on their occupation, judging the relative position of dispersive and localized states in these two compounds is very difficult. Charge densities for the localized states in the $\text{Nb}_{47}\text{O}_{116}$ and $\text{Nb}_{25}\text{O}_{62}$ are shown in Fig. 10. It seems very likely that both $\text{Nb}_{25}\text{O}_{62}$ and $\text{Nb}_{47}\text{O}_{116}$ possess only localized electrons, occupying a fraction of these localized states. Since $\text{Nb}_{47}\text{O}_{116}$ is a unit cell level intergrowth of $\text{Nb}_{22}\text{O}_{54}$ and $\text{Nb}_{25}\text{O}_{62}$, and $\text{Nb}_{22}\text{O}_{54}$ shows complete

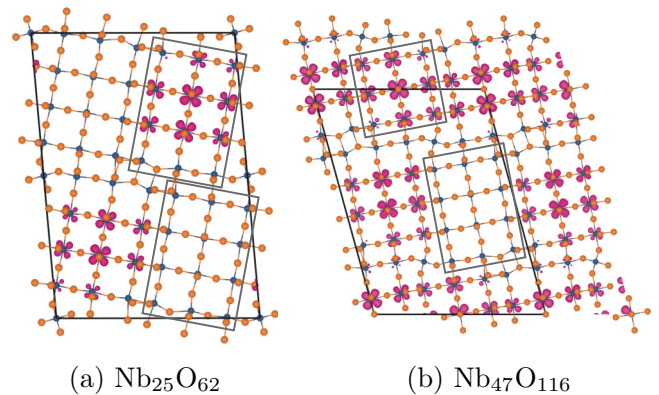


FIG. 10. Summed charge densities from bands in (a) $\text{Nb}_{25}\text{O}_{62}$ and (b) $\text{Nb}_{47}\text{O}_{116}$. Selected empty and filled localized states within blocks are framed by rectangles. The same (conventional) unit cell as in Fig. 1 is shown for $\text{Nb}_{25}\text{O}_{62}$, but a smaller primitive cell for $\text{Nb}_{47}\text{O}_{116}$.

localization of electrons, it is very likely that electrons should also fully localize in $\text{Nb}_{47}\text{O}_{116}$, at least in those parts of the structure that derive from $\text{Nb}_{22}\text{O}_{54}$.

D. $\text{H-Nb}_2\text{O}_5$

$\text{H-Nb}_2\text{O}_5$ is the high-temperature phase of niobium pentoxide, and crystallizes in space group $P2/m$ (Table I, Fig. 1). As the parent compound of the crystallographic shear structures, its electronic structure provides a reference. However, since it is fully oxidized, all niobium ions have a d^0 configuration and there are no electrons occupying the conduction band.

The band structure of $\text{H-Nb}_2\text{O}_5$ shows the presence of flat and dispersive bands (Fig. 11), similar to those found for the other shear phases above. However, the relative position of these bands depends on whether or not they are occupied. Doping by alkali metal ions is one way to introduce electrons into the conduction band, and in the particular case of $\text{H-Nb}_2\text{O}_5$ this has a practical relevance. Transition metal oxides in general, and the niobium-based oxide shear phases of this work, are used as electrodes within lithium-ion batteries. Like oxygen removal, lithium intercalation is a method to n -dope the material. Similar behavior can often be observed from charge doping and ion insertion, for example, in Na_xWO_3 [40]. The $\text{H-Nb}_2\text{O}_5$ phase has been studied extensively for lithium-ion battery applications and it is closely related to other shear phases that have been examined for the same purpose [19,41]. Inserting a single lithium per unit cell into the middle of the 3×4 block results in a localized state similar to those present in the niobium suboxides (Fig. 12). Note that the electron is entirely localized within the 3×4 block, with the 3×5 block remaining empty. Oxygen removal (as in the suboxides above) and lithium intercalation (examined here) clearly result in similar electronic structure features.

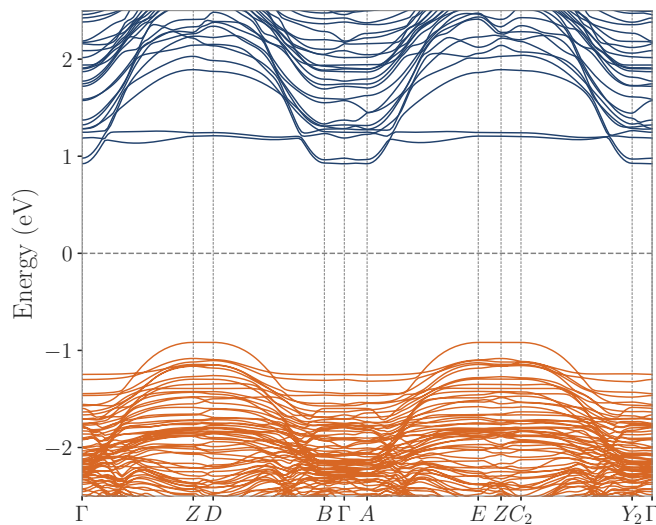


FIG. 11. Band structure of $\text{H-Nb}_2\text{O}_5$ (PBE). Valence and conduction bands are colored in orange and blue, respectively. Flat and dispersive bands are present above the Fermi level (dashed line) similar to those in $\text{Nb}_{22}\text{O}_{54}$ and $\text{Nb}_{12}\text{O}_{29}$, but are unoccupied.

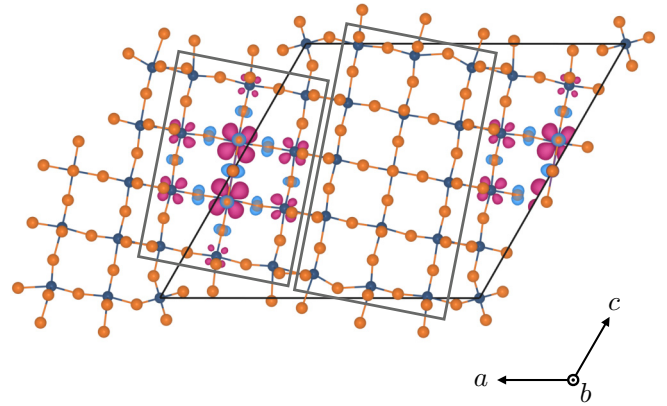


FIG. 12. Spin density plot of lithiated $\text{H-Nb}_2\text{O}_5$. A single lithium is located in the middle of the smaller block, inducing a localized state.

IV. DISCUSSION

Our results establish that the presence of defect-like flat bands and metallic conduction states is an innate feature of block-type structures. This coexistence arises due to the two different types of niobium sites present in the crystal structures; the central NbO_6 octahedra are purely corner shared, the distance between niobium atoms is larger and orbital overlap is reduced. This isolation results in localized electronic states, while along the crystallographic shear planes, where Nb-Nb distances are smaller and orbitals overlap more strongly, delocalized states are present. Each block can host one localized electron that is, rather unusually, spread over multiple niobium sites. This spread over multiple sites explains why single crystal x-ray diffraction studies on $\text{Nb}_{22}\text{O}_{54}$ and $o\text{-Nb}_{12}\text{O}_{29}$ do not show the presence of charge ordering [14], despite the detection of localized electrons by magnetic measurements [8]. As the electronic structure features are ultimately a result of the blocks as structural units, the same principles are likely to apply to other crystallographic shear phases in the $\text{WO}_3\text{-Nb}_2\text{O}_5$ and $\text{TiO}_2\text{-Nb}_2\text{O}_5$ phase diagrams.

Regarding possible electronic conduction mechanisms in the niobium suboxides, the relevant quantities are the filling fraction ν (number of e^- per block), and the energy gap between the flat and dispersive bands. For filling fractions of less than one, only localized states are filled ($\text{Nb}_{25}\text{O}_{62}$ and $\text{Nb}_{47}\text{O}_{116}$, Fig. 10) and electrons can hop from one filled block to another empty one by a polaron hopping mechanism. The hopping process will have an activation energy. However, in $\text{Nb}_{22}\text{O}_{54}$ all localized states are filled and this hopping mechanism becomes impossible. With no metallic conduction electrons present, thermal excitation from the defect-like flat bands into the dispersive conduction bands might provide the dominant mechanism, as illustrated in Fig. 13(a). This mechanism is reminiscent of doped semiconductors, and the activation energy associated with this process will depend on the separation between the flat and dispersive bands [cf. Figs. 5 and 13(a)]. Thermal excitation from flat into dispersive bands is also possible in $\text{Nb}_{25}\text{O}_{62}$ and $\text{Nb}_{47}\text{O}_{116}$, and could coexist with a polaron hopping mechanism. Finally, in the case of $\text{Nb}_{12}\text{O}_{29}$, all localized states are filled, but metallic

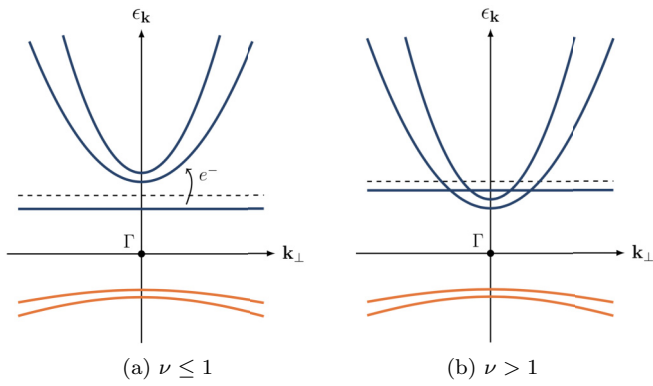


FIG. 13. Schematic of band structure for (a) filling fraction $\nu \leq 1 e^-/\text{block}$ and (b) $\nu > 1 e^-/\text{block}$. O $2p$ and Nb $4d$ dominated bands are colored in orange and blue, respectively. Fermi level is indicated by a dashed line, k_{\perp} designates reciprocal space vector associated with the real space direction perpendicular to the block plane. The relative position of flat and dispersive bands changes with the filling fraction ν .

conduction states are also partially filled [Fig. 13(b)]. The result is metallic conductivity, which does not require thermal activation. Both Rüscher *et al.* [12] and Cava *et al.* [8,10] have studied the conductivity of $\text{Nb}_2\text{O}_{5-\delta}$ compounds and observed that all phases except $\text{Nb}_{12}\text{O}_{29}$ show thermally activated conductivity. Rüscher *et al.* also noted the effectively one-dimensional electron transport properties along the block columns [13], which are consistent with the calculated band dispersions. In addition, the experiments of Cava *et al.* show semiconducting electronic behavior for $\text{Nb}_{25}\text{O}_{62}$ from 4–300 K; $\text{Nb}_{47}\text{O}_{116}$ and $\text{Nb}_{22}\text{O}_{54}$ exhibit semiconducting (thermally activated) conductivity from 0–250 K and from 0–100 K, respectively [8]. Beyond those temperatures, there is a metallic-like range of temperatures in which conductivity decreases again. Given this change in the temperature dependence of the conductivity from semiconducting to metallic, the flat bands associated with localized electrons

are likely shallow donor levels (Fig. 13). The complex interplay between electron localization and delocalization in the suboxides is more similar to phenomena occurring in semiconductors on n -type doping, but distinctly different from metal-insulator transitions in transition metal oxides [42]. Our results suggest that a similar phenomenon of crossover from localized to metallic conduction could occur on lithium doping of $\text{H-Nb}_2\text{O}_5$, which might be observed with electrochemical, spectroscopic, or magnetic measurements.

Magnetic susceptibility measurements on the suboxides show that the number of localized moments increases with the degree of reduction [8,9]. However, under the assumption that $g = 2$, the number of moments calculated from the measurements is smaller than the number of introduced electrons [8,9]. For $\text{Nb}_{12}\text{O}_{29}$, this is consistent with one delocalized and one localized electron, but for the remaining suboxides this apparent reduction in the number of local moments is unexpected. Our first-principles results are consistent with complete localization of electrons until a threshold is exceeded, and therefore all moments should be seen. Deviations in the g value of the electrons might explain the discrepancy. As the electrons in these suboxides are well localized, electron paramagnetic resonance (EPR) spectroscopy could provide some insight into the nature of the electronic states and the g values. For $\text{Nb}_{22}\text{O}_{54}$ in particular, the different shapes of the magnetic orbitals could be used to detect electrons occupying specific blocks. Another possibility is that localized electrons contributing magnetic moments coexist with magnetically inactive electrons in all suboxides, not just $\text{Nb}_{12}\text{O}_{29}$. However, we see no evidence for this in our calculations, and the thermally activated conductivity of $\text{Nb}_{22}\text{O}_{54}$, $\text{Nb}_{47}\text{O}_{116}$, and $\text{Nb}_{25}\text{O}_{62}$ seems inconsistent with the presence of magnetically inactive (Pauli-paramagnetic) metallic electrons.

Long-range antiferromagnetic order is observed only in the monoclinic $\text{Nb}_{12}\text{O}_{29}$ phase below 12 K, all other niobium suboxides are paramagnetic [8,10]. The Curie-Weiss constants of $\text{Nb}_2\text{O}_{5-\delta}$ are in the range of 0–24 K (0–2 meV), and indicate antiferromagnetic interactions that become stronger with increasing degree of reduction [8]. In $\text{Nb}_{25}\text{O}_{62}$ and

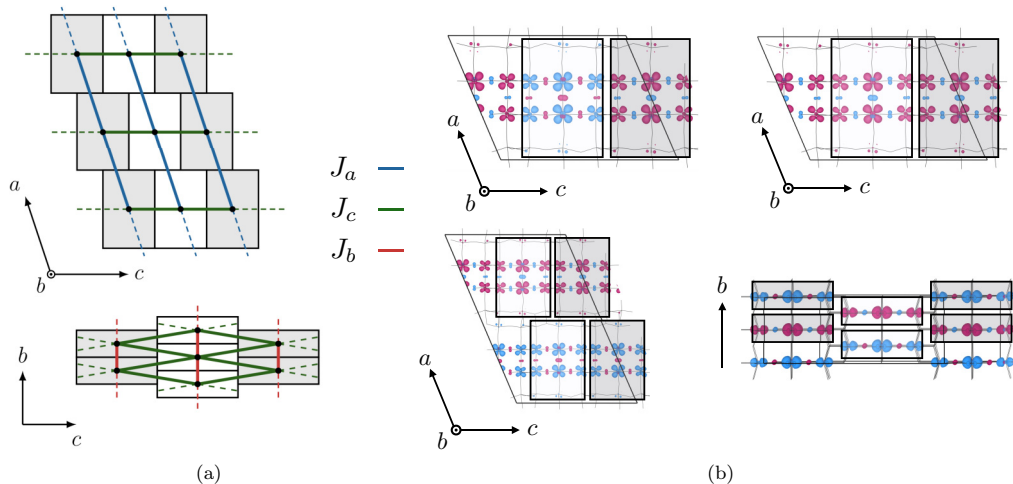


FIG. 14. (a) Possible spin-spin interactions J_i in monoclinic $\text{Nb}_{12}\text{O}_{29}$ along crystallographic a , b , and c directions. (b) Different spin arrangements in monoclinic $\text{Nb}_{12}\text{O}_{29}$. AFM along c (top left), fully FM (top right), AFM along a (bottom left), AFM along b (bottom right). White and gray blocks are offset by $0.5 b$ throughout.

$\text{Nb}_{47}\text{O}_{116}$, some of the localized states are empty (cf. Fig. 10), and the magnetic lattice is not fully filled. Independent of the strength of interaction, if not all spins have neighbors to couple with, or there is some randomness in the distribution of the spins, long-range magnetic order is unlikely to emerge. With first-principles calculations it is very difficult to address the question of why only $m\text{-Nb}_{12}\text{O}_{29}$ orders, but $\text{Nb}_{22}\text{O}_{54}$ and $o\text{-Nb}_{12}\text{O}_{29}$ do not, since the energy differences between different magnetic states are very small. However, we can discuss the possible spin-spin interactions simply based on the shape and orientation of the magnetic orbital within the crystal structure. We will focus in particular on monoclinic $\text{Nb}_{12}\text{O}_{29}$, but similar considerations apply to the other suboxides. The magnetic orbital lies within the plane of the block. The two closest distances (two neighbors) between spins (two nearest neighbors) are along the block columns, with a separation of about 3.8 \AA [J_b , Fig. 14(a) bottom]. By symmetry, the interaction with spins in the four next-nearest neighboring blocks along c , that are offset by $0.5 b$, has to be the same [J_c , distance 10.6 \AA , Fig. 14(a) bottom]. In addition to that, each block is connected to two blocks on the same level in monoclinic $\text{Nb}_{12}\text{O}_{29}$ along the a direction [J_a , distance 15.9 \AA , Fig. 14(a) top], and four others offset by $0.5 b$ along a [distance 15.3 \AA , Fig. 14(a) top]. Different spin arrangements are easily obtained from DFT calculations [spin densities are shown in Fig. 14(b)], but the energy differences between them are very small (few meV), and change significantly with the level of theory (PBE or PBE + U). Energy differences of a few meV are consistent with the interaction strengths obtained experimentally. The lowest energy magnetic ordering found in our calculations is antiferromagnetic along the c direction [Fig. 14(b), top left].

V. CONCLUSION

We have shown that the electronic structure features common to n -doped crystallographic shear phases include (1) effectively one-dimensional flat and dispersive bands corresponding to localized and delocalized electronic states, (2) electron localization in orbitals spanning the block planes,

and (3) the partition of localized and delocalized states between central and peripheral niobium sites. Structural block units are also present in $\text{WO}_3\text{-Nb}_2\text{O}_5$ [43] and $\text{TiO}_2\text{-Nb}_2\text{O}_5$ [44,45] phases, and many of these mixed-metal shear phases have been explored as lithium-ion battery electrodes [19,46]. The principles laid out in this work are likely transferable to these compounds, and are important for the interpretation of spectroscopic and electrochemical data.

The niobium suboxides show a transition from localized to delocalized electrons, but it is much smoother than commonly observed for metal-insulator transitions in transition metal oxides. In fact, our results portray the suboxides to be closer to n -doped semiconductors, but with a limited capacity for localized electrons. Once a filling threshold is exceeded, delocalized metallic electrons are simply added to existing localized electrons. This process is likely to occur in heavily lithium-doped shear phases during battery operation. Similarly, the experimentally observed crossover from localized to delocalized electronic behavior in WO_{3-x} [47] might have the same underlying mechanism, as WO_{3-x} phases also exhibit some amount of crystallographic shear. More broadly, the niobium suboxides are an elegant example of the interplay between crystal and electronic structure, and the balance between electron localization and delocalization in oxides of an early transition metal.

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- [1] K. Kato and S. Tamura, *Acta Crystallogr. Sect. B* **31**, 673 (1975).
 - [2] K. Kato, *Acta Crystallogr. Sect. B* **32**, 764 (1976).
 - [3] H. Schäfer, R. Gruehn, and F. Schulte, *Angew. Chem., Int. Ed. Engl.* **5**, 40 (1966).
 - [4] R. S. Roth and A. D. Wadsley, *Acta Crystallogr.* **19**, 26 (1965).
 - [5] R. S. Roth and A. D. Wadsley, *Acta Crystallogr.* **19**, 42 (1965).
 - [6] M. Kunz and I. D. Brown, *J. Solid State Chem.* **115**, 395 (1995).
 - [7] I. Bersuker, *The Jahn-Teller Effect* (Cambridge University Press, Cambridge, 2006).
 - [8] R. J. Cava, B. Batlogg, J. J. Krajewski, H. F. Poulsen, P. Gammel, W. F. Peck, and L. W. Rupp, *Phys. Rev. B* **44**, 6973 (1991).
 - [9] C. H. Ruscher and M. Nygren, *J. Phys.: Condens. Matter* **3**, 3997 (1991).
 - [10] R. J. Cava, B. Batlogg, J. J. Krajewski, P. Gammel, H. F. Poulsen, W. F. Peck Jr., and L. W. Rupp Jr., *Nature (London)* **350**, 598 (1991).
 - [11] E. N. Andersen, T. Klimczuk, V. L. Miller, H. W. Zandbergen, and R. J. Cava, *Phys. Rev. B* **72**, 033413 (2005).
 - [12] C. Ruscher, E. Salje, and A. Hussain, *J. Phys. C* **21**, 3737 (1988).
 - [13] C. Ruscher, *Physica C: Superconductivity* **200**, 129 (1992).
 - [14] T. McQueen, Q. Xu, E. Andersen, H. Zandbergen, and R. Cava, *J. Solid State Chem.* **180**, 2864 (2007).
 - [15] A. Lappas, J. E. L. Waldron, M. A. Green, and K. Prassides, *Phys. Rev. B* **65**, 134405 (2002).
 - [16] J.-G. Cheng, J.-S. Zhou, J. B. Goodenough, H. D. Zhou, C. R. Wiebe, T. Takami, and T. Fujii, *Phys. Rev. B* **80**, 134428 (2009).
 - [17] C. M. Fang, M. A. van Huis, Q. Xu, R. J. Cava, and H. W. Zandbergen, *J. Mater. Chem. C* **3**, 651 (2014).

- [18] K.-W. Lee and W. E. Pickett, *Phys. Rev. B* **91**, 195152 (2015).
- [19] K. J. Griffith, K. M. Wiaderek, G. Cibir, L. E. Marbella, and C. P. Grey, *Nature (London)* **559**, 556 (2018).
- [20] R. Li, Y. Qin, X. Liu, L. Yang, C. Lin, R. Xia, S. Lin, Y. Chen, and J. Li, *Electrochim. Acta* **266**, 202 (2018).
- [21] J. E. L. Waldron, M. A. Green, and D. A. Neumann, *J. Phys. Chem. Solids* (A Special Issue on the Occasion of the 65th Birthday of Professor Peter Day FRS) **65**, 79 (2004).
- [22] S. J. Clark, M. D. Segall, C. J. Pickard, P. J. Hasnip, M. I. J. Probert, K. Refson, and M. C. Payne, *Z. Kristallogr.* **220**, 567 (2005).
- [23] D. R. Hamann, M. Schlüter, and C. Chiang, *Phys. Rev. Lett.* **43**, 1494 (1979).
- [24] D. Vanderbilt, *Phys. Rev. B* **41**, 7892 (1990).
- [25] M. Hellenbrandt, *Crystallogr. Rev.* **10**, 17 (2004).
- [26] J. P. Perdew, K. Burke, and M. Ernzerhof, *Phys. Rev. Lett.* **77**, 3865 (1996).
- [27] H. J. Monkhorst and J. D. Pack, *Phys. Rev. B* **13**, 5188 (1976).
- [28] See Supplemental Material at <http://link.aps.org/supplemental/10.1103/PhysRevB.99.075151> for crystallographic information files (CIF) of the structures dealt with in this article.
- [29] S. L. Dudarev, G. A. Botton, S. Y. Savrasov, C. J. Humphreys, and A. P. Sutton, *Phys. Rev. B* **57**, 1505 (1998).
- [30] H. H. Nahm and C. H. Park, *Phys. Rev. B* **78**, 184108 (2008).
- [31] J. Heyd, G. E. Scuseria, and M. Ernzerhof, *J. Chem. Phys.* **118**, 8207 (2003).
- [32] Y. Hinuma, G. Pizzi, Y. Kumagai, F. Oba, and I. Tanaka, *Comput. Mater. Sci.* **128**, 140 (2017).
- [33] A. Togo and I. Tanaka, [arXiv:1808.01590](https://arxiv.org/abs/1808.01590).
- [34] A. J. Morris, R. J. Nicholls, C. J. Pickard, and J. R. Yates, *Comput. Phys. Commun.* **185**, 1477 (2014).
- [35] C. J. Pickard and M. C. Payne, *Phys. Rev. B* **59**, 4685 (1999).
- [36] C. J. Pickard and M. C. Payne, *Phys. Rev. B* **62**, 4383 (2000).
- [37] M. J. Rutter, *Comput. Phys. Commun.* **225**, 174 (2018).
- [38] K. Momma and F. Izumi, *J. Appl. Crystallogr.* **44**, 1272 (2011).
- [39] M. Evans, ML-*evs/matador*, <https://bitbucket.org/ml-evs/matador>.
- [40] A. D. Walkingshaw, N. A. Spaldin, and E. Artacho, *Phys. Rev. B* **70**, 165110 (2004).
- [41] K. J. Griffith, A. C. Forse, J. M. Griffin, and C. P. Grey, *J. Am. Chem. Soc.* **138**, 8888 (2016).
- [42] M. Imada, A. Fujimori, and Y. Tokura, *Rev. Mod. Phys.* **70**, 1039 (1998).
- [43] R. S. Roth and A. D. Wadsley, *Acta Crystallogr.* **18**, 724 (1965).
- [44] A. D. Wadsley, *Acta Crystallogr.* **14**, 660 (1961).
- [45] A. D. Wadsley, *Acta Crystallogr.* **14**, 664 (1961).
- [46] B. Guo, X. Yu, X.-G. Sun, M. Chi, Z.-A. Qiao, J. Liu, Y.-S. Hu, X.-Q. Yang, J. B. Goodenough, and S. Dai, *Energy Environ. Sci.* **7**, 2220 (2014).
- [47] E. Salje and B. Güttler, *Philos. Mag. B* **50**, 607 (1984).